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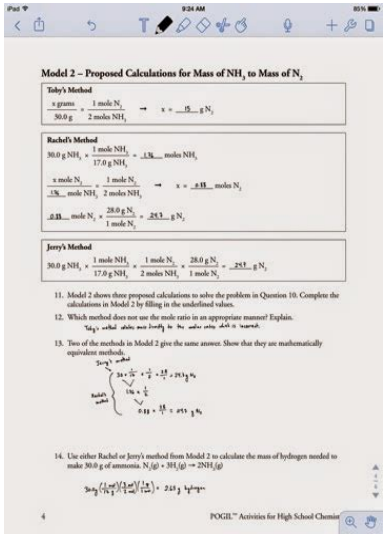

Model 2 – Proposed Calculations for Mass of NH_3 to Mass of N_2

Toby's Method

$\frac{x \text{ grams}}{30.0 \text{ g}} = \frac{1 \text{ mole } N_2}{2 \text{ moles } NH_3} \rightarrow x = 15 \text{ g } N_2$

Rachel's Method

$30.0 \text{ g } NH_3 \times \frac{1 \text{ mole } NH_3}{17.0 \text{ g } NH_3} = 1.76 \text{ moles } NH_3$

$\frac{x \text{ mole } N_2}{1.76 \text{ mole } NH_3} = \frac{1 \text{ mole } N_2}{2 \text{ moles } NH_3} \rightarrow x = 0.88 \text{ moles } N_2$

$0.88 \text{ mole } N_2 \times \frac{28.0 \text{ g } N_2}{1 \text{ mole } N_2} = 24.5 \text{ g } N_2$

Jerry's Method

$30.0 \text{ g } NH_3 \times \frac{1 \text{ mole } NH_3}{17.0 \text{ g } NH_3} \times \frac{1 \text{ mole } N_2}{2 \text{ moles } NH_3} \times \frac{28.0 \text{ g } N_2}{1 \text{ mole } N_2} = 24.7 \text{ g } N_2$

11. Model 2 shows three proposed calculations to solve the problem in Question 10. Complete the calculations in Model 2 by filling in the underlined values.

12. Which method does not use the mole ratio in an appropriate manner? Explain.

13. Two of the methods in Model 2 give the same answer. Show that they are mathematically equivalent methods.

14. Use either Rachel or Jerry's method from Model 2 to calculate the mass of hydrogen needed to make 30.0 g of ammonia. $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$

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